

## **Engineering Chemistry Theory (CH110/CH210)**

**2017-18**

**Branch – CV, CTM AND ENV**

### **Course Objective**

To impart the knowledge of applied aspects of chemistry and utilizing the same for the technological advancement in various discipline of engineering.

### **Course outcome**

On successful completion of this course, the students will be able to

- CO-1. Understand the concepts of electrochemistry and electrochemical phenomenon involved in the energy storage and energy conversion devices
- CO-2. Able to know the mechanism of corrosion and its control and application of electrochemical concepts for surface modification techniques
- CO-3. Know the importance of petroleum products as conventional sources of energy and the importance of polymers as an engineering material, their synthesis and application in automobile, electronic and aerospace applications
- CO-4. Know the importance of analysis of water and waste water and the technological applications of nanomaterials.
- CO-5. Able to understand the different approaches to types of chemical bonding and give an account of VB theory that predicts the geometry of molecules in terms of concepts of hybridisations. To know the metallurgical process in iron and steel. Understand the colloidal system and interfacial phenomena.

**JSS Mahavidyapeetha**  
**JSS Science and Technology University**  
**Sri Jayachamarajendra College of Engineering, Mysore-570 006**

**Department :Chemistry**

**Subject: Engineering Chemistry**  
**Branch: CV, CTM & ENV**

**Syllabus forthe Academic year 2017-2018**

**Sub code: CH 110/210**

**Credits: 4**

**Total hours: 52 hrs**

**UNIT –I**

**ELECTROCHEMISTRY**

Introduction, Single electrode potential – definition & origin, Significance of Nernst equation. Electrodes – Types-Reference electrodes – calomel electrode & Ag/AgCl electrode. Measurement of emf using calomel electrode. Concentration cells – definition, construction and working. Numerical problems. Ion selective electrode – glass electrode, determination of pH using glass electrode. Numerical problems.

**BATTERY TECHNOLOGY**

Introduction, definition, battery characteristics, classification – primary, secondary and reserve with examples. Modern batteries-construction, working and applications of Nickel-Metal hydride, Nickel-Cadmium, Lithium-MnO<sub>2</sub>, Li-ion and Zn-air batteries.

**FUEL CELLS**

Introduction, Classification, Construction, working and applications of H<sub>2</sub>-O<sub>2</sub> and methanol-oxygen fuel cells.

**11hrs**

**UNIT-II**

**CORROSION SCIENCE**

Corrosion- definition, types-chemical and electrochemical corrosion. Electrochemical theory of corrosion, Factors affecting the rate of corrosion-nature of metal, nature of corrosion product, relative areas of anode and cathode, temperature and pH. Types of corrosion – differential metal corrosion, differential aeration corrosion (pitting and waterline corrosion), stress corrosion-caustic embrittlement in boilers.

Corrosion control – Inorganic coatings-anodizing and phosphating, metal coating- galvanizing and tinning. Corrosion inhibitors-cathodic and anodic. Cathodic protection- sacrificial anode and impressed current techniques, Anodic protection.

## **ELECTROPLATING AND ELECTROLESS PLATING**

Importance, significance of polarization, decomposition potential and over-voltage in electroplating processes. Electroplating process: Effects of variables on the nature of electro deposit – current density, metal salt and electrolyte concentration, metal ion concentration, temperature, pH of the bath, additives – brighteners, levelers, structure modifier and wetting agents, throwing power of the bath. Surface preparation – by using solvents, alkali, acid and electropolishing, Electroplating of Cr and Ni.

**Electroless plating** – Differences between electroplating and electroless plating, advantages of electroless plating, Electroless plating of copper on PCB.

**12hrs**

## **UNIT –III**

### **ENERGY SOURCES**

**Chemical Fuels-** Definition, classification; calorific value-definition, Gross and Net calorific values (SI units). Determination of calorific value of a solid/liquid fuel using Bomb calorimeter, numerical problems, Petroleum cracking-fluidized catalytic cracking. Reformation of petrol. Knocking and its mechanism, octane number, cetane number, prevention of knocking, anti-knocking agents, unleaded petrol.

### **HIGH POLYMERS**

Polymers - Introduction, Glass transition temperature (T<sub>g</sub>) – parameters affecting T<sub>g</sub> and significance of T<sub>g</sub>. Synthesis, properties and applications of PTFE, PMMA and PU

**Elastomers** – definition, deficiencies of natural rubber, advantages of synthetic rubber, Vulcanisation of rubber. Synthesis and applications of neoprene and butyl rubber.

**Adhesives** – definition, synthesis, properties and applications of epoxy resin.

**Conducting polymers** – definition and mechanism of conduction in polyacetylene.

**10 hrs**

## **UNIT-IV**

### **WATER CHEMISTRY**

Introduction, hardness - types, units of hardness. Determination of hardness of water by EDTA method. Water analysis – estimation of chloride, fluoride and nitrate. Determination of DO by winkler's method. Bio-chemical oxygen Demand and Chemical Oxygen Demand. Numerical problems on BOD and COD. Desalination of water – electro dialysis and reverse osmosis.

### **CHEMISTRY OF NANOMATERIALS**

Introduction, Definition, classification of nanomaterials-0D spheres and clusters, 1D Nano fibres, wires and rods (multilayers), 2D films, plates and networks, (Ultrafine-grained overlayers), 3D nanomaterials. General properties of nanomaterials, Synthesis of nanomaterials – top down and bottom up approach-methods –sol gel method and chemical vapour deposition method.Applications of nanomaterials and nanotechnology.

**10hrs**

## UNIT-V

### CHEMICAL BONDING AND THEORY OF IONIZATION

Ionisation energy, electron affinity and periodicity. Conditions and energetics of formation of ionic and covalent bonds. Covalent bonds – polarity of covalent compounds and characteristics of covalent compounds. Difference between ionic and covalent compounds. Partial ionic characters of covalent bonds. Valence Bond approach, Sigma and Pi bonds. Geometry of molecules and hybridization. Metallic bonds.

### METALLURGY

Principles of Metallurgy – Introduction, Characteristics of metals, Physical properties of metals, Sources of elements, occurrence of metals, slags and fluxes. Metals and metallic character, non-metals, metalloids. Classification of ores, Furnaces, Ore dressing.

Iron and steel- occurrence, impurities in ores and their effect, commercial forms of iron. Allotropic forms of iron. Properties of iron, effect of impurities on properties of cast iron.

### COLLOIDAL STATE

Colloids – Introduction, Particle size, Classification of Colloids, Difference between colloidal solution, true solution and suspension. Preparation of colloids, Purification of colloidal solutions, Properties of colloids, stabilization of colloidal solutions, Interaction between hydrophobic and hydrophilic colloids, Brownian movement, Emulsions, Gels. Coagulation and Flocculation. Applications of colloids.

9hrs

### Reference Books:

1. A text book of Engineering Chemistry by Jain and Jain, Dhanapatrai Publications, New Delhi.
2. Engineering Chemistry by Uppal, Khanna Publishers, Sixth Edition, 2001.
3. Principles of Physical Chemistry by B.R. Puri, L.R.Sharma& M.S. Pathania, S.Nagin Chand and Co., 33<sup>rd</sup> Ed., 1992.
4. A text book of Physical Chemistry by P.L.Soni and O.P.Dharma.
5. A text book of Polymer science by V.R. Gowarikar& others New-age publications.
6. Corrosion Engineering-by M. G. Fontana, McGraw Hill Publications.
7. Text book of Polymer science by F.W.Billmeyer, John, Wiley and Sons, 1994.
8. Environmental chemistry by Stanley E. Manahan, 7<sup>th</sup> edition, Lewis publishers, 2000.
9. Hand book of Nanotechnology, BharathBhushan, Spinger-Verlag Berlin Heidelberg New York.2004.
10. A text book of Engineering Chemistry by S.S. Dara, S. Chand & Company Ltd., New Delhi.
11. A text book of Engineering Chemistry by R. V. Gadag and A. Nityananda Shetty, Published by I. K. International Publishing House Pvt. Ltd. New Delhi.
12. A text book of Engineering Chemistry by B. S. Jai Prakash, R. Venugopal, Shivakumaraiah, PushpaIyengar



## **Engineering Chemistry Laboratory CH11L/CH22L**

### **Course objective**

To impart practical knowledge in the field of chemical analysis with respect to conventional and instrumental methods of chemical analysis.

### **Course outcome**

- CO-1. Assessment of water quality parameter like hardness & alkalinity and judging the suitability of water for domestic and industrial applications.
- CO-2. Understanding the importance of alloys as an engineering material and applying to analyzing them by instrumental techniques and assessment of quality of haematite used for metallurgical operation.
- CO-3. Assessment of waste water quality parameters like BOD and COD in determining the extent of pollution.
- CO-4. Applications of analytical techniques like potentiometry, pH metry and conductometry for accurate chemical analysis.
- CO-5. Applications of analytical techniques like colorimetry for accurate chemical analysis.

**JSS Mahavidyapeetha**  
**JSS Science and Technology University**  
**Sri Jayachamarajendra College of Engineering, Mysore-570 006**

**Department : Chemistry**

**Subject: Engineering Chemistry**  
**Branch: CV, CTM & ENV**

**Syllabus for the Academic year 2017-2018**

**Sub code: CHL 110/210**

**Credits: 1.5**

**Lab hours: 3hrs**

**PART – A**

1. Determination of total hardness of water by EDTA method.
2. Determination of total alkalinity of water sample and identification of the type and extent of alkalinity.
3. Determination of percentage of Copper in brass in the given sample of brass by iodometric method.
4. Determination of Iron in haematite solution using standard solution of potassium dichromate by external indicator method.
5. Determination of dissolved oxygen of the given water sample by Winkler's method
6. Determination of Chemical Oxygen Demand of an industrial effluent.

**PART – B**

1. Determination of  $p^{K_a}$  of weak acid using pH meter.
2. Estimation of iron in stainless steel/ FAS potentiometrically using standard solution of potassium dichromate.
3. Conductometric estimation of HCl and  $CH_3COOH$  present in a mixture using standard solution of sodium hydroxide.
4. Determination of copper colorimetrically using ammonia as the complexing agent.
5. Determination of iron (III) by colorimetric method using potassium thiocyanate as the complexing agent.
6. Determination of equivalent conductance of strong electrolyte at infinite dilution.
7. Flame photometric method of determining sodium in water sample.

**Reference Books:**

1. Vogels text book of quantitative inorganic analysis, revised by J.Bassett, R.C.Denny, G.H.Jeffery, 4<sup>th</sup> Ed.
2. Applied chemistry theory and practice by O. P. Vermani and A. K. Narula, second edition.
3. Water and waste water analysis by American  $\alpha$ -method (APHS).