

Engineering Chemistry Theory (CH110/CH210)

2017-18

Branch – EC, EE & IT

Course Objective

To impart the knowledge of applied aspects of chemistry and utilizing the same for the technological advancement in various discipline of engineering.

Course outcome

On successful completion of this course, the students will be able to

- CO-1. Understand the concepts of electrochemistry and electrochemical phenomenon involved in the energy storage and energy conversion devices.
- CO-2. Able to know the mechanism of corrosion and its control and application of electrochemical concepts for surface modification techniques.
- CO-3. Know the importance of petroleum products as conventional sources of energy and the importance of polymers as an engineering material, their synthesis and application in automobile, electronic and aerospace applications.
- CO-4. Know the importance of analysis of water and waste water and the technological applications of nanomaterials.
- CO-5. Able to understand the concepts of supercapacitors and their applications.
To understand the construction and working of PV cells with their applications.
To know the importance of liquid crystal in different applications.

JSS Mahavidyapeetha
JSS Science and Technology University
Sri Jayachamarajendra College of Engineering, Mysore-570 006

Department: Chemistry

Subject: Engineering Chemistry
Branch: EC, EE & IT

Syllabus for the Academic year 2017-2018

Sub code: CH 110/210

Credits: 4

Total hours: 52 hrs

UNIT –I

ELECTROCHEMISTRY

Introduction, Single electrode potential – definition & origin, Significance of Nernst equation. Electrodes – Types-Reference electrodes – calomel electrode & Ag/AgCl electrode. Measurement of emf using calomel electrode. Concentration cells – definition, construction and working. Numerical problems. Ion selective electrode – glass electrode, determination of pH using glass electrode. Numerical problems.

BATTERY TECHNOLOGY

Introduction, definition, battery characteristics, classification – primary, secondary and reserve with examples. Modern batteries-construction, working and applications of Nickel-Metal hydride, Nickel-Cadmium, Lithium-MnO₂, Li-ion and Zn-air batteries.

FUEL CELLS

Introduction, Classification, Construction, working and applications of H₂-O₂ and methanol-oxygen fuel cells.

11hrs

UNIT-II

CORROSION SCIENCE

Corrosion- definition, types-chemical and electrochemical corrosion. Electrochemical theory of corrosion, Factors affecting the rate of corrosion-nature of metal, nature of corrosion product, relative areas of anode and cathode, temperature and pH. Types of corrosion – differential metal corrosion, differential aeration corrosion (pitting and waterline corrosion), stress corrosion-caustic embrittlement in boilers.

Corrosion control – Inorganic coatings-anodizing and phosphating, metal coating- galvanizing and tinning. Corrosion inhibitors-cathodic and anodic. Cathodic protection- sacrificial anode and impressed current techniques, Anodic protection.

ELECTROPLATING AND ELECTROLESS PLATING

Importance, significance of polarisation, decomposition potential and over-voltage in electroplating processes. Electroplating process: Effects of variables on the nature of electro deposit – current density, metal salt and electrolyte concentration, metal ion concentration, temperature, pH of the bath, additives – brighteners, levelers, structure modifier and wetting agents, throwing power of the bath. Surface preparation – by using solvents, alkali, acid and electropolishing, Electroplating of Cr and Ni.

Electroless plating – Differences between electroplating and electroless plating, advantages of electroless plating, Electroless plating of copper on PCB.

12hrs

UNIT –III

ENERGY SOURCES

Chemical Fuels- Definition, classification; calorific value-definition, Gross and Net calorific values (SI units). Determination of calorific value of a solid/liquid fuel using Bomb calorimeter, numerical problems, Petroleum cracking-fluidized catalytic cracking. Reformation of petrol. Knocking and its mechanism, octane number, cetane number, prevention of knocking, anti-knocking agents, unleaded petrol.

HIGH POLYMERS

Polymers - Introduction, Glass transition temperature (T_g) – parameters affecting T_g and significance of T_g. Synthesis, properties and applications of PTFE, PMMA and PU

Elastomers – definition, deficiencies of natural rubber, advantages of synthetic rubber, Vulcanisation of rubber. Synthesis and applications of neoprene and butyl rubber.

Adhesives– definition, synthesis, properties and applications of epoxy resin.

Conducting polymers – definition and mechanism of conduction in polyacetylene.

10hrs

UNIT-IV

WATER CHEMISTRY

Introduction, hardness - types, units of hardness. Determination of hardness of water by EDTA method. Water analysis – estimation of chloride, fluoride and nitrate. Determination of DO by winkler's method. Bio-chemical oxygen Demand and Chemical Oxygen Demand. Numerical problems on BOD and COD. Desalination of water – electro dialysis and reverse osmosis.

CHEMISTRY OF NANOMATERIALS

Introduction, Definition, classification of nanomaterials-0D spheres and clusters, 1D Nano fibres, wires and rods (multilayers), 2D films, plates and networks, (Ultrafine-grained overlayers), 3D nanomaterials. General properties of nanomaterials, Synthesis of nanomaterials – top down and bottom up approach-methods –sol gel method and chemical vapour deposition method.Applications of nanomaterials and nanotechnology.

10hrs

UNIT-V

SUPER CAPACITORS

Introduction, Types of supercapacitors: Electrochemical Double-Layer Capacitors-activated carbons, carbon aerogels and nanotubes. Pseudocapacitors-conducting polymers and metal oxides. Advantages, Limitations and Applications.

RENEWABLE SOURCES OF ENERGY

Introduction, Importance of solar (PV) cells, Photovoltaic cell, the p/n junction of a solar cell. Working of solar cell. Doping of silicon for photovoltaic cells and chemical properties for solar grade silicon. Production of solar grade silicon from quartz. Dye sensitized solar cells.

LIQUID CRYSTALS

Introduction, Liquid crystal phase, Classification: Thermotropic liquid crystal and Lyotropic liquid crystals, Chemical constitution and Liquid crystalline behavior, homologous series, Applications- LCD. LED, OLED & Plasma display.

9hrs

Reference Books:

1. A text book of Engineering Chemistry by Jain and Jain, Dhanapatrai Publications, New Delhi.
2. Engineering Chemistry by Uppal, Khanna Publishers, Sixth Edition, 2001.
3. Principles of Physical Chemistry by B.R. Puri, L.R.Sharma& M.S. Pathania, S.Nagin Chand and Co., 33rd Ed., 1992.
4. A text book of Physical Chemistry by P.L.Soni and O.P.Dharma.
5. A text book of Polymer science by V.R. Gowarikar& others New-age publications.
6. Corrosion Engineering-by M. G. Fontana, McGraw Hill Publications.
7. Text book of Polymer science by F.W.Billmeyer, John, Wiley and Sons, 1994.
8. Environmental chemistry by Stanley E. Manahan, 7th edition, Lewis publishers, 2000.
9. Hand book of Nanotechnology, BharathBhushan, Spinger-Verlag Berlin Heidelberg New York.2004.
10. A text book of Engineering Chemistry by S.S. Dara, S. Chand & Company Ltd., New Delhi.
11. A text book of Engineering Chemistry by R. V. Gadag and A. Nityananda Shetty, Published by I. K. International Publishing House Pvt. Ltd. New Delhi.
12. A text book of Engineering Chemistry by B. S. Jai Prakash, R. Venugopal, Shivakumaraiah, PushpaIyengar

Engineering Chemistry Laboratory CH11L/CH22L

Course objective

To impart practical knowledge in the field of chemical analysis with respect to conventional and instrumental methods of chemical analysis.

Course outcome

- CO-1. Assessment of water quality parameter like hardness & alkalinity and judging the suitability of water for domestic and industrial applications.
- CO-2. Understanding the importance of alloys as an engineering material and applying to analyzing them by instrumental techniques and assessment of quality of haematite used for metallurgical operation.
- CO-3. Assessment of waste water quality parameters like BOD and COD in determining the extent of pollution.
- CO-4. Applications of analytical techniques like potentiometry, pH metry and conductometry for accurate chemical analysis.
- CO-5. Applications of analytical techniques like colorimetry for accurate chemical analysis.

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Syllabus for the Academic year 2017-2018

Sub code: CHL 110/210

Credits: 1.5

Lab hours: 3hrs

PART – A

1. Determination of total hardness of water by EDTA method.
2. Determination of total alkalinity of water sample and identification of the type and extent of alkalinity.
3. Determination of percentage of Copper in brass in the given sample of brass by iodometric method.
4. Determination of Iron in haematite solution using standard solution of potassium dichromate by external indicator method.
5. Determination of dissolved oxygen of the given water sample by Winkler's method
6. Determination of Chemical Oxygen Demand of an industrial effluent.

PART – B

1. Determination of p^{K_a} of weak acid using pH meter.
2. Estimation of iron in stainless steel/ FAS potentiometrically using standard solution of potassium dichromate.
3. Conductometric estimation of HCl and CH_3COOH present in a mixture using standard solution of sodium hydroxide.
4. Determination of copper colorimetrically using ammonia as the complexing agent.
5. Determination of iron (III) by colorimetric method using potassium thiocyanate as the complexing agent.
6. Determination of equivalent conductance of strong electrolyte at infinite dilution.
7. Flame photometric method of determining sodium in water sample.

Reference Books:

1. Vogels text book of quantitative inorganic analysis, revised by J.Bassett, R.C.Denny, G.H.Jeffery, 4th Ed.
2. Applied chemistry theory and practice by O. P. Vermani and A. K. Narula, second edition.
3. Water and waste water analysis by American α -method (APHS).